



Standard Practices for Preparation of Sample Containers and for Preservation of Organic Constituents¹

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1. Scope

1.1 These practices cover the various means of (1) preparing sample containers used for collection of waters to be analyzed for organic constituents and (2) preservation of such samples from the time of sample collection until the time of analysis.

1.2 The sample preservation practice is dependent upon the specific analysis to be conducted. See Section 9 for preservation practices listed with the corresponding applicable general and specific constituent test method. The preservation method for waterborne oils is given in Practice D3325. Use of the information given herein will make it possible to choose the minimum number of sample preservation practices necessary to ensure the integrity of a sample designated for multiple analysis. For further considerations of sample preservation, see the *Manual on Water*.²

1.3 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.* For specific hazard statements, see 6.7, 6.24, and 8.1.3.

2. Referenced Documents

2.1 ASTM Standards:³

D1129 Terminology Relating to Water

D1193 Specification for Reagent Water

D1252 Test Methods for Chemical Oxygen Demand (Dichromate Oxygen Demand) of Water

D1783 Test Methods for Phenolic Compounds in Water
D2036 Test Methods for Cyanides in Water
D2330 Test Method for Methylene Blue Active Substances (Withdrawn 2011)⁴
D2579 Test Method for Total Organic Carbon in Water (Withdrawn 2002)⁴
D2580 Test Method for Phenols in Water by Gas-Liquid Chromatography
D2908 Practice for Measuring Volatile Organic Matter in Water by Aqueous-Injection Gas Chromatography
D3113 Test Methods for Sodium Salts of EDTA in Water (Withdrawn 2005)⁴
D3325 Practice for Preservation of Waterborne Oil Samples
D3371 Test Method for Nitriles in Aqueous Solution by Gas-Liquid Chromatography (Withdrawn 2002)⁴
D3534 Test Method for Polychlorinated Biphenyls (PCBs) in Water (Withdrawn 2003)⁴
D3590 Test Methods for Total Kjeldahl Nitrogen in Water
D3695 Test Method for Volatile Alcohols in Water by Direct Aqueous-Injection Gas Chromatography
D3856 Guide for Management Systems in Laboratories Engaged in Analysis of Water
D3871 Test Method for Purgeable Organic Compounds in Water Using Headspace Sampling
D3921 Test Method for Oil and Grease and Petroleum Hydrocarbons in Water
D3973 Test Method for Low-Molecular Weight Halogenated Hydrocarbons in Water
D4129 Test Method for Total and Organic Carbon in Water by High Temperature Oxidation and by Coulometric Detection
D4165 Test Method for Cyanogen Chloride in Water
D4193 Test Method for Thiocyanate in Water
D4281 Test Method for Oil and Grease (Fluorocarbon Extractable Substances) by Gravimetric Determination
D4282 Test Method for Determination of Free Cyanide in Water and Wastewater by Microdiffusion
D4374 Test Methods for Cyanides in Water—Automated Methods for Total Cyanide, Weak Acid Dissociable Cyanide, and Thiocyanate

¹ These practices are under the jurisdiction of ASTM Committee D19 on Water and are the direct responsibilities of Subcommittee D19.06 on Methods for Analysis for Organic Substances in Water.

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² *Manual on Water*, ASTM STP 442, ASTM, 1969.

³ For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

⁴ The last approved version of this historical standard is referenced on www.astm.org.

- D4515** Practice for Estimation of Holding Time for Water Samples Containing Organic Constituents (Withdrawn 2006)⁴
- D4657** Test Method for Polynuclear Aromatic Hydrocarbons in Water (Withdrawn 2005)⁴
- D4744** Test Method for Organic Halides in Water by Carbon Adsorption/Microcoulometric Detection (Withdrawn 2002)⁴
- D4763** Practice for Identification of Chemicals in Water by Fluorescence Spectroscopy
- D4779** Test Method for Total, Organic, and Inorganic Carbon in High Purity Water by Ultraviolet (UV) or Persulfate Oxidation, or Both, and Infrared Detection (Withdrawn 2002)⁴
- D4839** Test Method for Total Carbon and Organic Carbon in Water by Ultraviolet, or Persulfate Oxidation, or Both, and Infrared Detection
- D4841** Practice for Estimation of Holding Time for Water Samples Containing Organic and Inorganic Constituents
- D4983** Test Method for Cyclohexylamine, Morpholine, and Diethylaminoethanol in Water and Condensed Steam by Direct Aqueous Injection Gas Chromatography (Withdrawn 2002)⁴
- D5175** Test Method for Organohalide Pesticides and Polychlorinated Biphenyls in Water by Microextraction and Gas Chromatography
- D5176** Test Method for Total Chemically Bound Nitrogen in Water by Pyrolysis and Chemiluminescence Detection
- D5315** Test Method for Determination of N-Methyl-Carbamoyloximes and N-Methylcarbamates in Water by Direct Aqueous Injection HPLC with Post-Column Derivatization
- D5316** Test Method for 1,2-Dibromoethane and 1,2-Dibromo-3-Chloropropane in Water by Microextraction and Gas Chromatography
- D5317** Test Method for Determination of Chlorinated Organic Acid Compounds in Water by Gas Chromatography with an Electron Capture Detector
- D5412** Test Method for Quantification of Complex Polycyclic Aromatic Hydrocarbon Mixtures or Petroleum Oils in Water
- D5475** Test Method for Nitrogen- and Phosphorus-Containing Pesticides in Water by Gas Chromatography with a Nitrogen-Phosphorus Detector (Withdrawn 2011)⁴
- D5790** Test Method for Measurement of Purgeable Organic Compounds in Water by Capillary Column Gas Chromatography/Mass Spectrometry
- D5812** Test Method for Determination of Organochlorine Pesticides in Water by Capillary Column Gas Chromatography (Withdrawn 2011)⁴

3. Terminology

3.1 *Definitions*—For definitions of terms used in this practice, refer to Terminology **D1129**.

4. Significance and Use

4.1 There are four basic steps necessary to obtain meaningful analytical data: preparation of the sample container, sampling, sample preservation, and analysis. In fact these four

basic steps comprise the analytical method and for this reason no step should be overlooked. Although the significance of preservation is dependent upon the time between sampling and the analysis, unless the analysis is accomplished within 2 h after sampling, preservation is preferred and usually required.

5. Apparatus

5.1 *Forced Draft Oven*, capable of operating at 275 to 325°C.

5.2 *Sample Bottle*, borosilicate or flint glass.

NOTE 1—High density polyethylene (HDPE) bottles and caps have been demonstrated to be of sufficient quality to be compatible for all tests except pesticides, herbicides, polychlorinated biphenyls, and volatile organics. However, this bottle cannot be recycled.

5.3 *Sample Bottle Cap*, TFE-fluorocarbon or aluminum foil-lined.

NOTE 2—Even these liners have some disadvantages. TFE is known to collect some organic constituents, for example, PCBs. Aluminum foil will react with samples that are strongly acid or alkaline. Clean TFE liners as described in 7.1. Replace aluminum foil with new foil after each use.

5.4 *Sample Vial*, glass.

5.5 *Septa*, PTFE-faced with screw cap lid and matching aluminum foil disks.

6. Reagents and Materials

6.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society.⁵ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

6.2 *Purity of Water*—Unless otherwise indicated, reference to water shall be understood to mean reagent water conforming to Specification **D1193**, Type II and demonstrated to be free of specific interference for the test being performed.

6.3 *Acetic Acid Buffer Solution* (pH 4)—Dissolve 6.0 g of sodium acetate in 75 mL of water. Add 30 mL of glacial acetic acid, with stirring.

6.4 *Acetone*.

6.5 *Acid Buffer Solution* (pH 3.75)—Dissolve 125 g of potassium chloride and 70 g of sodium acetate trihydrate in 500 mL of water. Add 300 mL of glacial acetic acid and dilute to 1 L.

6.6 *Ascorbic Acid*.

6.7 *Chromic Acid Cleaning Solution*—To a 2-L beaker, add 35 mL of saturated sodium dichromate solution followed by 1 L of sulfuric acid (sp gr 1.84) with stirring. (**Warning**—Use rubber gloves, safety goggles, and protective clothing when

⁵ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmaceutical Convention, Inc. (USPC), Rockville, MD.

preparing and handling this corrosive cleaning agent that is a powerful oxidant. Store the reagent in a glass bottle with a glass stopper.)

6.8 *Detergent*, formulated for cleaning laboratory glassware.

6.9 *Hydrochloric Acid*—Concentrated HCl (sp gr 1.19).

6.10 *Hydrochloric Acid* (1 + 2)—To 200 mL of water, carefully add 100 mL of hydrochloric acid (see 6.9). Store in a glass-stoppered reagent bottle.

6.11 *Ice*, crushed wet.

6.12 *Lead Acetate Test Paper*.

6.13 *Lead Acetate Solution*—Dissolve 50 g of lead acetate in water and dilute to 1 L.

6.14 *Lead Carbonate*, powdered.

6.15 *Lime, Hydrated*, powdered.

6.16 *Mercuric Chloride*.

6.17 *Monochloroacetic Acid Buffer* (pH 3)—Prepare by mixing 156 mL of chloroacetic acid solution (236.2 g/L) and 100 mL of potassium acetate solution (245.4 g/L).

6.18 *Nitric Acid*—Concentrated HNO₃ (sp gr 1.42).

6.19 *Phosphate Buffer*—Dissolve 138 g of sodium dihydrogen phosphate in water and dilute to 1 L. Refrigerate this solution.

6.20 *Phosphate Solution*—Dissolve 33.8 g of potassium dihydrogen phosphate in 250 mL of water.

6.21 *Phosphoric Acid*—Concentrated H₃PO₄ (sp gr 1.83).

6.22 *Phosphoric Acid Solution* (1 + 1)—Dilute 1 vol of phosphoric acid (sp gr 1.83).

6.23 *pH Paper*, narrow range for pH < 2, pH > 12, and pH 5 to 7.

6.24 *Potassium Iodide–Starch Test Paper*.

6.25 *Sodium Bisulfate*.

6.26 *Sodium Bisulfite Solution*—Dissolve 2 g of sodium bisulfite in 1 L of water and adjust to pH 2 by the slow addition of H₂SO₄ (1 + 1). (**Warning**—Prepare and use this reagent in a well ventilated hood to avoid exposure to SO₂ fumes.)

6.27 *Sodium Sulfite Solution* (0.1 M)—Transfer approximately 10.3 g of sodium sulfite to a 1-L volumetric flask. Dilute to volume with water.

6.28 *Sodium Thiosulfate*.

6.29 *Sodium Hydroxide Pellets*.

6.30 *Mercuric Chloride* (10 mg/mL)—Dissolve 100 mg of HgCl₂ in reagent water and dilute to 10 mL.

6.31 *Sulfuric Acid* (1 + 1)—Slowly and carefully add 1 vol of sulfuric acid (see 6.27) to 1 vol of water, stirring and cooling the solution during addition.

7. Preparation of HDPE Sample Bottles

7.1 Wash the bottles with two 100-mL portions of HCl (1 + 2) and rinse with three 100-mL portions of water. These

volumes of wash and rinse portions are recommended for 1-L sample bottles; therefore, use proportionate volumes for washing and rinsing sample bottles of a different volume.

8. Preparation of Glass Sample Bottles and Vials

8.1 *Solvent-Detergent/Chromic Acid Preparation of Glass Sample Bottles*:

8.1.1 Rinse the container with 100 mL of dilute detergent or acetone. For some residues, a few alternative detergent and acetone rinses may be more satisfactory. Then rinse at least three times with tap water followed by a reagent water rinse to remove the residual detergent or acetone, or both.

8.1.2 Rinse the container with 100 mL of chromic acid solution, returning the chromic acid to its original container after use. Then rinse with at least three 100-mL portions of tap water followed by a reagent water rinse.

8.1.3 Rinse the container with 100 mL of NaHSO₃ solution to remove residual hexavalent chromium. (**Warning**—Carry out this step in a hood to prevent exposure to SO₂ fumes.)

8.1.4 Rinse the container with water until sulfuric acid and its vapors have been removed. Test rinsings for acid with a pH meter or an appropriate narrow range pH paper. Rinsings should have a pH approximately the same as the water used for rinsing.

8.1.5 When the last trace of NaHSO₃ has been removed, wash with three additional 100-mL portions of water. Allow to drain. This procedure is for 1-L sample containers, therefore, use proportionate volumes for washing and rinsing sample containers of a different volume.

8.1.6 Heat for a minimum of 4 h (mouth up) in a forced draft oven at 275 to 325°C. Upon cooling, fit the bottles with caps and the vials with septa.

NOTE 3—For some tests, heating may not be required. Refer to the individual method to determine the necessity for this treatment.

8.2 Machine Washing Glass Sample Bottles and Vials:

NOTE 4—Machine washing of narrow mouth sample bottles may not yield acceptable results.

8.2.1 Rinse the container with 100 mL of chromic acid solution, returning the chromic acid to its original container after use. Then rinse with at least three 100-mL portions of tap water.

8.2.2 Machine wash in accordance with the machine manufacturer's instructions using a detergent and 90°C water.

8.2.3 Remove the bottles from the machine and rinse them with two 100-mL portions of HCl (1 + 2), followed with three 100-mL portions of water.

8.2.4 Heat for a minimum of 4 h (mouth up) in a forced draft oven at 275 to 325°C. Upon cooling, fit the bottles with caps and the vials with septa (see Note 3).

9. Sample Preservation

9.1 Depending upon the type of analysis required, use any one or a combination of the following methods of sample preservation (see Tables 1-3, Annex A1, and Annex A2).

9.1.1 Adjust the pH. An adjustment to neutral pH is usually prescribed when chemical reactions, such as hydrolysis, are to be avoided. Adjustment to an extreme pH, for example, <2, is